

Model Question Paper

Chemistry

Class IX going to X

Time – 22.5 minutes

F.M. $15 \times 4 = 60$

For every correct answer examinee will be awarded +4 marks and for every wrong answer it will be – 1 mark

- SI unit of volume is
(A) cm^3 (B) dm^3 (C) m^3 (D) None of these
- Today's temperature is 2°C lesser than that of yesterday. Then the change of temperature in Kelvin scale in
(A) 2K (B) 275K (C) 271K (D) None of these
- Dry ice is
(A) Solid H_2O (B) Solid CO_2 (C) Solid SO_2 (D) Solid NO_2
- If a gas of density $\frac{1}{100000}$ of density of normal air is super cooled we get
(A) Plasma stable of matter (B) Bose-Einstein condensate
(C) Super cooled liquid (D) None of these
- Diameter of the solute particles in a solution should be
(A) Less than 1 nm (B) More than 100 nm
(C) Less than 100 nm but more than 1 nm (D) None of these
- 40 gm of common salt are there in 320 gm of water in a solution. Calculate the concentration in terms of mass by mass percentage
(A) 1.11% (B) 11.1% (C) 10.1% (D) 10.2%
- Which of the followings is homogeneous mixture
(A) Sol (B) Solution (C) Suspension (D) Both (1) & (2)

8. A colloidal solution where in both dispersed phase and dispersing medium are liquid is called
 (A) Sol (B) Aerosol (C) Foam (D) Emulsion
9. A colour has been made by adding three pigments. These three pigments can be separated by
 (A) Fractional distillation (B) Distillation
 (C) Using separating funnel (D) Chromatography
10. Solubility of which of the following substances decreases with increase in temperature.
 (A) Potassium chloride (B) Potassium nitrate
 (C) Ammonium chloride (D) Sodium chloride
11. For the reactions $A + B = C + D$. The law of conservation of mass states
 (A) Mass of A = Mass of C & Mass of B = Mass of D
 (B) Mass of A = Mass of D & Mass of B = Mass of C
 (C) Mass of A + Mass of B = Mass of C + Mass of D
 (D) All of the above
12. Match the column
- | Elements | Atomic mass |
|-------------|-------------|
| A – Carbon | P – 32 |
| B – Sodium | Q – 40 |
| C – Calcium | R – 12 |
| D – Sulphur | S – 23 |
- (A) A – P (B) A – Q (C) A – R (D) A – R
 B – S B – R B – P B – S
 C – Q C – S C – S C – Q
 D – R D – P D – Q D – P
13. Which about the law of constant proportion is incorrect
 (A) It was given by Proust
 (B) It states that a compound is made of two or more definite elements
 (C) It deals with the cases where two element combine to give more than one compounds
 (D) The mass ratio of two or more elements combining to give a compound is constant
14. A is an element with atomic number 15 and B has atomic number 17. The binary compound made by these two elements is
 (A) AB_3 (B) AB_5 (C) Both (D) None

15. An element has three natural isotopes with atomic masses X, Y and Z. They are present in nature in a ratio A : B : C. The average atomic mass of this element is

(A) $\frac{AX + BY}{A + B} + \frac{BY + CZ}{B + C} + \frac{AX + CZ}{A + C}$

(B) $\frac{AX + BY + CZ}{A + B + C}$

(C) $\frac{X + Y + Z}{3}$

(D) $\frac{X + Y + Z}{A + B + C}$

